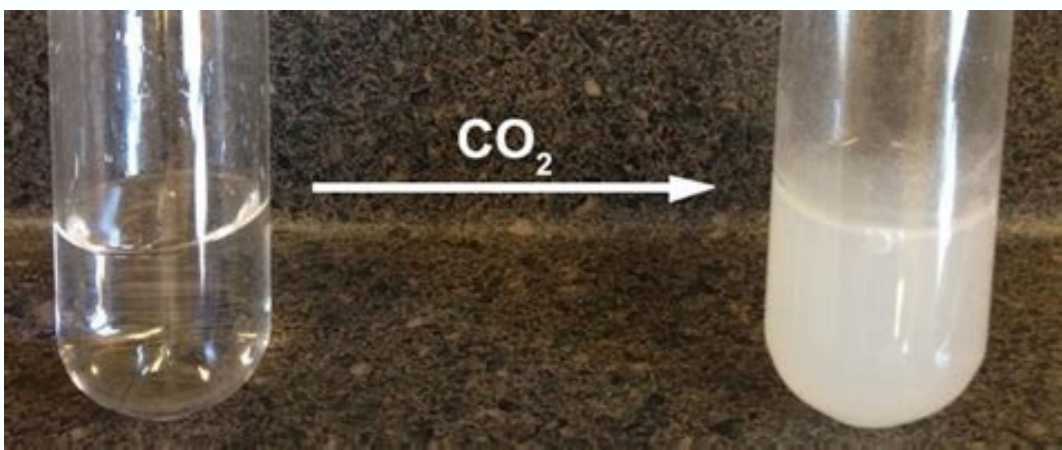
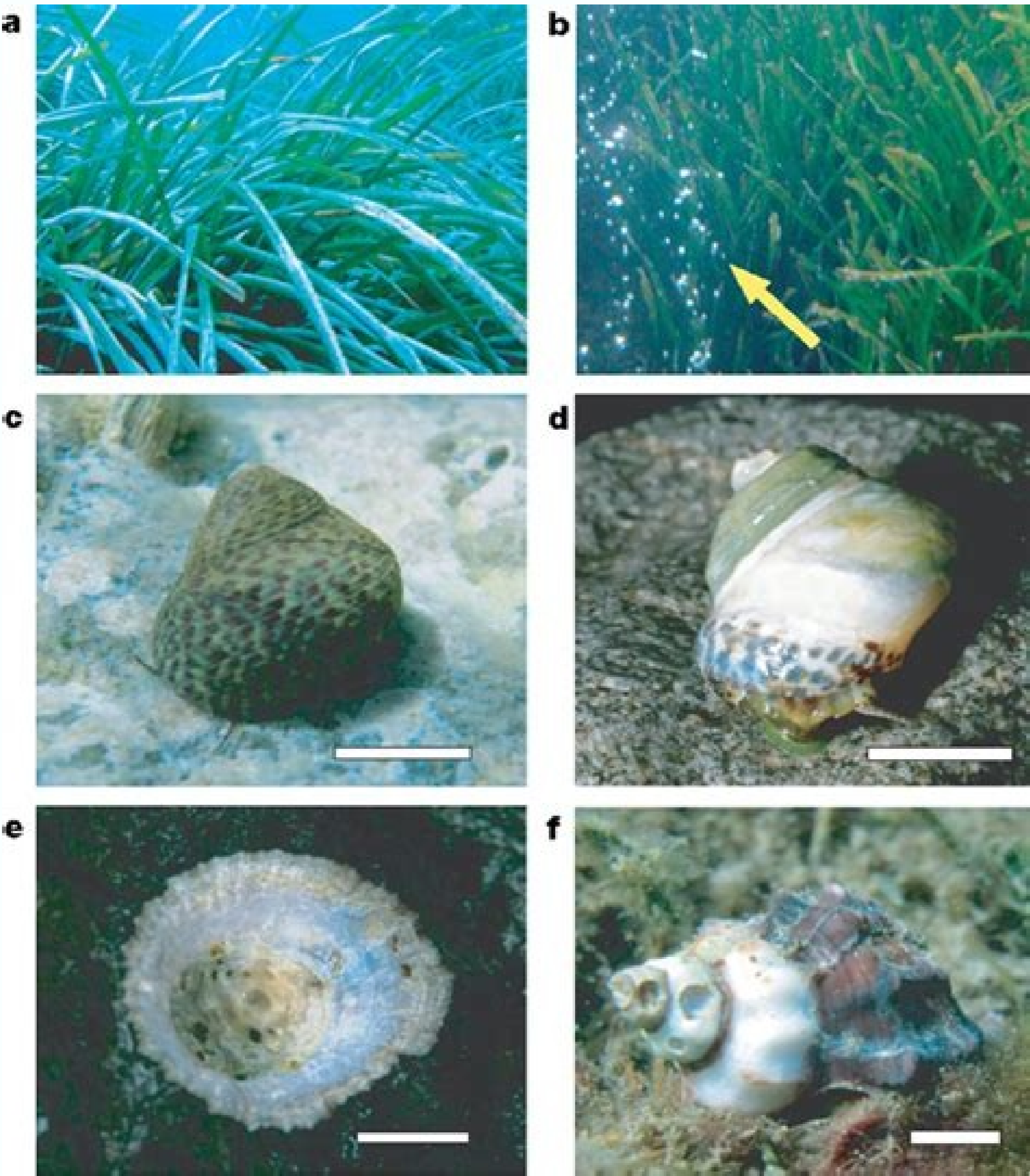


Calcium carbonate and water react to form

I'm not robot!



Calcium carbonate and water reaction equation.

By Shayna, Facty StaffUpdated: Nov 10, 2021Calcium carbonate and calcium phosphate are effective tools for adding extra calcium to one's diet. Both can be accessed over the counter, are readily available, and can be found in multiple formats such as pills, chewable tablets, and powders. While both are effective supplements, they have benefits and drawbacks them should be considered before adding them to a daily vitamin regimen. While the best way to maintain a healthy level of calcium is to eat a balanced diet, supplements can aid in ensuring continued health and preventing certain diseases. Calcium phosphate is a dietary supplement used to treat calcium deficiency. Specifically, it is naturally occurring calcium that has bonded to phosphate. For dietary purposes, it is available in tablet, powder, and chewable forms. It typically contains an impressive 39% calcium, while other supplements have only 20%. An alternative to calcium phosphate, calcium carbonate — the latter is a mineral in limestone — has the added benefits of an antacid to neutralize stomach acid. It is one of the most common options, as it dissolves well, for rapid absorption. Calcium carbonate is available in both capsule and powder form. Typically, it is 40% calcium by weight. Calcium supplements such as calcium phosphate and calcium carbonate provide several benefits. While the primary use is to build and maintain strong bones and prevent osteoporosis, that is not all they can achieve. Combining calcium supplements with vitamin D (which aids in the absorption of calcium) has been found to protect against cancer, diabetes, and high blood pressure. Every person, young and old, should be meeting a daily minimum requirement of calcium. Averages change based on age and sex. For those who are not reaching this amount through natural sources such as dairy and leafy greens, a doctor may recommend a calcium supplement. In particular, the following groups may need calcium supplementation: VegansLactose intolerant peopleAnyone receiving long-term corticosteroid treatmentThose with a diet high in protein or sodiumThose with bowel or digestive issues While any calcium delivery system has its advantages, phosphate is required for optimal bone health, making calcium phosphate ideal for those with a deficiency, typically associated with celiac disease, kidney issues, alcohol disorder, and those who take too many antacids. Another benefit of calcium phosphate is that the amount of calcium is higher than most other supplements. While phosphate is essential for bone health, too much can lead to osteoporosis or issues with kidney function. Because most people get sufficient phosphates through diet, adding more via calcium supplements may cause issues. Because of this, calcium phosphate may only be appropriate for those with other health issues. Be sure to discuss any supplements or medications with a doctor before adding them to a daily plan. Calcium carbonate can be used in the treatment of many disorders including dyspepsia, hungry bone syndrome, hypocalcemia, and tumor lysis syndrome. Similar to calcium phosphate, it has a high volume of calcium, making it ideal for bone deficiencies. It can also be used to decrease the levels of phosphates in blood. Additionally, it is the cheapest calcium supplement on the market and can double as an antacid. While the benefits of calcium carbonate are numerous, it is important to note that not all advantages have been empirically proven. Additionally, it requires vitamin D to be effective. A doctor can confirm what calcium supplement is best for each patient's needs. Both calcium phosphate and calcium carbonate can effectively increase calcium in the body, containing higher percentages of calcium than many other supplements. That said, both can react negatively with other medications. Regardless of the supplement, it is important to discuss it with a doctor prior to use.Both supplements are available in different forms for ease of use and are absorbed more effectively with vitamin D. Supplements are good for people who have limitations to their diet that prevent them getting enough of a vital nutrient, or those with deficiencies preventing their bodies from properly absorbing a nutrient. However, natural sources are usually the best option, and in the case of calcium, there are many choices. Excellent natural sources of calcium — that could remove the need for a supplement — include MilkGreens like okra and kaleBone-in fish, such as sardines and canned salmonSoy products fortified with calcium Navigation Bar MAIN Demos Calcium is a silvery-white metal; it is relatively soft, but much harder than sodium metal. Calcium is a member of the alkaline-earth metals (Group II on the periodic table); these metals react vigorously with water, although not as violently as the Group I metals such as sodium or potassium: $\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ In the following demonstration, a chunk of calcium metal is dropped into a beaker of distilled water. After a second or so, the calcium metal begins to bubble vigorously as it reacts with the water, producing hydrogen gas, and a cloudy white precipitate of calcium hydroxide. The presence of the hydroxide is demonstrated by the addition of a few drops of phenolphthalein indicator, which turns the solution pink, indicating that the solution is basic. Video Clip: REAL, 4.31 MB !!! Hazards !!! Hydrogen gas is produced during the course of this reaction. If you are not collecting the gas, perform the procedure in a fume hood or a well-ventilated area to allow the gas to dissipate. Procedures Producing Hydrogen Gas from Calcium Metal: Lee R. Summerlin, Christie L. Borgford, and Julie B. Ealy, Chemical Demonstrations: A Sourcebook for Teachers, Volume 2, 2nd ed. Washington, D.C.: American Chemical Society, 1988, p. 51-52. References John Emsley, The Elements, 3rd ed. Oxford, Clarendon Press, 1998, p. 48-49. David L. Heiserman, Exploring Chemical Elements and their Compounds. New York: TAB Books, 1992, p. 84-87. Martha Windholz (ed.), The Merck Index, 10th ed. Rahway: Merck & Co., Inc., 1983. Chemical compound Calcium carbonate Names IUPAC name Calcium carbonate Other names Aragonite; calcite; chalk; lime; limestone; marble; oyster; pearl Identifiers CAS Number 471-34-1 Y 3D model (JSmol) Interactive imageChEBI CHEBI:3311 Y ChEMBL ChEMBL1200539 N ChemSpider 9708 Y DrugBank DB06724 ECHA InfoCard 100.006.765 EC Number 207-439-9 E number E170 (colours) KEGG D00932 Y PubChem CID 10112 RTECS number FF9335000 UNII H0G9379FGK Y CompTox Dashboard (EPA) DTXSID3036238 InChI InChI=1S/CH2O3.Ca/c2-1(3)4;/h(H2,2,3,4);/q;+2/p-2 YKey: VTYLEPIZMXCLO-UHFFFAOYSA-L YInChI=1/CH2O3.Ca/c2-1(3)4;/h(H2,2,3,4);/q;+2/p-2Key: VTYLEPIZMXCLO-NUQVWONBAS SMILES [Ca+2].[O-]C([O-])=OC(=O)([O-])[O-].[Ca+2] Properties Chemical formula CaCO3 Molar mass 100.0869 g/mol Appearance Fine white powder; chalky taste Odor odorless Density 2.711 g/cm3 (calcite)2.83 g/cm3 (aragonite) Melting point 1,339 °C (2,442 °F; 1,612 K) (calcite) 825 °C (1,517 °F; 1,098 K) (aragonite)[4][5] Boiling point decomposes Solubility in water 0.013 g/L (25 °C)[1][2] Solubility product (Ksp) 3.3×10−9[3] Solubility in dilute acids soluble Acidity (pKa) 9.0 Magnetic susceptibility (χ) −3.82×10−5 cm3/mol Refractive index (nD) 1.59 Structure Crystal structure Trigonal Space group 32/m Thermochemistry Std molarentropy (So298) 93 J·mol−1·K−1[6] Std enthalpy of formation ($\Delta\text{H}^\ominus_{298}$) −1207 kJ·mol−1[6] Pharmacology ATC code A02AC01 (WHO) A12AA04 (WHO) Hazards NFPA 704 (fire diamond) 0 0 0 Lethal dose or concentration (LD, LC): LD50 (median dose) 6450 mg/kg (oral, rat) NIOSH (US health exposure limits): PEL (Permissible) TWA 15 mg/m3 (total) TWA 5 mg/m3 (resp)[7] Safety data sheet (SDS) ICSC 1193 Related compounds Other anions Calcium bicarbonate Other cations Beryllium carbonateMagnesium carbonateStrontium carbonateBarium carbonateRadium carbonate Related compounds Calcium sulfate Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa). N verify (what is "N"?) Infobox references Chemical compound Crystal structure of calcite Calcium carbonate is a chemical compound with the formula CaCO3. It is a common substance found in rocks as the minerals calcite and aragonite (most notably as limestone, which is a type of sedimentary rock consisting mainly of calcite) and is the main component of eggshells, gastropod shells, shellfish skeletons and pearls. Calcium carbonate is the active ingredient in agricultural lime and is created when calcium ions in hard water react with carbonate ions to create limescale. It has medical use as a calcium supplement or as an antacid, but excessive consumption can be hazardous and cause hypercalcemia and digestive issues.[8] Chemistry Calcium carbonate shares the typical properties of other carbonates. Notably it reacts with acids, releasing carbon dioxide (technically speaking, carbonic acid, but that disintegrates quickly to CO2 and H2O): $\text{CaCO}_3\text{(s)} + 2\text{H}^+\text{(aq)} \rightarrow \text{Ca}^{2+}\text{(aq)} + \text{CO}_2\text{(g)} + \text{H}_2\text{O(l)}$ releases carbon dioxide upon heating, called a thermal decomposition

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